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Sociale Menu Learning Link or Chile is an initiative of the Degree Department of Chile University. Resources to learn Biency and mathematical sciences in this section will find content of biology, chemistry, physical and mathematics, oriented towards the activation of Previous knowledge and the resolution of frequent doubts content of the first university. CID-base solutions, pH and limit a possible even if there are different definitions of acids and bases, in this section we will describe the foundations of killing acids and bases. At the end of the 1800s, Swedish scientist disappeared Arrhenius proposed that water can dissolve many compounds to separate them in their individual ions and defined acids and bases. Lid and bases according to Arrhenius: they are compounds containing hydrogen and can be dissolved in water releasing hydrogen ions to the solution. For example, hydrochloric acid (HCl) is dissolved in water as follows: bases: compounds that are dissolved in water by releasing oxide ions (Oh-) to the solution. For example, a typical base is sodium hydroxide (NaOH): in 1923, the Johannes scientist brÄnnsted and English Thomas Lowry published, in an independent form, jobs that the Rerhenius redefined theory. Lid and bases according to BÄÄÄnnsted and Lowry could mention: any compound that can give a hydrogen ion. Acids are commonly known as proton donors, because a hydrogen ion is simply a protonÄ (H +). Bases: any compound that can accept a hydrogen ion. The definition of BÄÄÄnnsted-Lowry manages to explain why the substances that do not contain Oh- can act as a basis. pH: Ä, in the definition of BrÄÄnnsted-Lowry, acids and bases are linked to the concentration of the hydrogen ion present. Acids increase the concentration of hydrogen ions, Ä, while the basics decrease to the concentration of hydrogen ions (when accepting them). As a result, the acidity or alkalinity of a solution can be measured by its concentration of hydrogen ions. The number of a solution number of acidity is obtained by applying the logarithm mathematical function to the concentration of a solution protons, which is called pH, Ä and generated a pH scale of solutions less than 1 m protons, this expression is " The following: here's how you can calculate the pH of different solutions. The pH, in practical terms, measures the degree of acidity and alkalinity of aqueous solutions. In 1909, the biochemistry danger SÄÄÄren was invented pH scale (from 1 to 14) to measure acidity. Strong acids are completely ionized in aqueous solution: on the other hand, those who dissociate very little in aqueous solution, such as acid acid, the dissociation reaction is in balance: the compounds that have the property of behaving like acids or bases They are called amphoters or ampholites (H2O). Packaging solutions are solutions formed by acid acids or bases and their salt. They have the ability to dampen or balance pH variations that occur in the solution when adding acids or bases over it. For example, in the blood shocked main absorption systems: H2CO3 / HCO3- and H2PO4- / hPO4 = When adding weak acid has a certain concentration of a strong bod base of concentration parity, a title curve is obtained with The following features: In the figure you see an area where the pH is practically not varied (buffer zone) Ä and corresponding to the Pka (-Log Ka) of the drop drop. The concentration of acid without neutralizing and salt formed at this point are the same, the pH at this point corresponds to the Pka (loga log). The structure uses to prepare the solutions between a o Basic Debil and its salt in the concentration range that allows you to absorb quantities of acid or base without producing large variations of pH. PH of a shock absorber: for a pillow consisting of Ch3-CooH and Ch3-Coona, the acid is partially dissociated and the salt is completely dissociated the pH of a cushioning solution according to the concentrations of acid and its salt can be calculated Using Henderson-Hasselbach equation: in this way, this type of solution has the property to keep the pH stable in an interval, this is the case of the conjugated carbon-ion bicarbonate couple, which in the blood "as a buffer And keeps the pH at 7.4 +/- 0.5. Informing the following pages will be able to find activities and exercises to improve your abilities and strategies for these contents: maybe it may be interested, we don't find the games with your profile . If desired, you can look for specific resources for keyword or according to your career! An acid-base reaction reaction or neutralization is a chemical reaction that takes place between an acid, for example chloridr acid ICO and a base, sodium hydroxide, production rooms, sodium chloride or salt .. The term "salt" describes any ion compound whose cut comes from a base (Na + del Naoh) and whose attachment derives from an acid (HCL CL-). Neutralization reactions are generally exotic, which means they emanate energy in the form of heat. They are usually called neutralization because by reacting a panic with a base, which neutralize their respective properties. There are several concepts that provide alternative definitions for reaction mechanisms involved in these reactions, and their application in dissolution problems connected to them. The word neutralization can be interpreted as a cancellation or elimination, which is not very distant to reality. When an acid is mixed with a base of both species react in different degrees that depend largely on the concentrations and volumes of acid and the base as an illustrative can be considered the reaction of a strong acid that mixes with a weak base, it will be Completely neutralized, while a strong acid portion will remain in solution, depending on the moles that reacted with the base. There can be three additional alternatives that arise from the mixture of a rock with a base: a strong acid is mixed with a strong base: when this happens, the species that will be in dissolution will be the one that is higher than the other . A weak acid is mixed with a strong base: dissolution will be essential, since it will be the base that remains in the reaction. A weak acid is mixed with a weak base: if this happens, the acidity of a solution depends on the constant acid acidity of acid acid and concentrations is the base and the Ä cido. Despite the differences in definitions, its importance is revealed in the different methods of analysis, when applied to acid-base reactions of gaseous or liquid species, or when the car. Cido or base can be something less evident. The first of these scientific concepts of acids and bases were provided by the father of the chemistry, the frank Ä © s Antoine Lavoisier, around 1776. [1] Ä c Ä, -

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